Chapter 10 – Review for Test

1. Calculate the formula mass of copper I oxide.
2. Calculate the formula mass of potassium permanganate.
3. What is the percent carbon in Acetic Acid?
4. What is the percent barium in barium fluoride?
5. Determine the empirical formula and name for a compound where the sample contains 32.8 % Cr and 67.2 % Cl.
6. Determine the empirical formula and name for a compound where the sample contains 8.29 % Al, 32.7 % Cl. and 59.0 % O.

1. Determine the molecular formula for a compound that has a molar mass of 78 g/mol and is 92.3 % C and 7.70 % H.
2. Determine the molecular formula for a compound that has a molar mass of 695 g/mol is 26.7 % P, 12.1 % N, and 61.2 % Cl.
3. What is the formula for a hydrate of **strontium oxalate**: mass of hydrate is 100.00 grams, the mass of the anhydrous salt is 90.70 g.
4. What is the density of fluorine gas at STP?
5. What is the molar mass of a gas that has a density of 3.57 g/L at STP?